

**Table 8.1: Fundamental physical formats, constants, and relationships between the units of the elementary quantities count, amount, and charge.**

Format	Name	Definitions	Numerical value	Unit *
$\underline{N}$	count, molecular #	count $N_X = N \cdot U_X$	$\triangle$ MU = x	
$\underline{n}$	amount, molar	amount $n_X = N_X \cdot N_A^{-1}$	$\square$ MU = mol	
$\underline{q}$	charge, electrical	charge $Q_{eX} = n_X \cdot z_X \cdot F$	$\circledast$ MU = C	
$\underline{N}$	Boltzmann constant §	$k = f \cdot e = R / N_A = 1.380\,649 \cdot 10^{-23}$	$\text{J} \cdot \text{x}^{-1} \cdot \text{K}^{-1}$	
$\underline{n}$	gas constant	$R = f \cdot F = k \cdot N_A = 8.314\,462\,618$	$\text{J} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$	
$\underline{q}$	electromotive constant §	$f = k / e = R / F = 8.617\,333\,262 \cdot 10^{-5}$	$\text{J} \cdot \text{C}^{-1} \cdot \text{K}^{-1}$	
$\underline{N} / \underline{n}$	Avogadro constant §	$N_A = R / k = F / e = 6.022\,140\,76 \cdot 10^{23}$	$\text{x} \cdot \text{mol}^{-1}$	
$\underline{q} / \underline{N}$	elementary charge §,¶	$e = k / f = F / N_A = 1.602\,176\,634 \cdot 10^{-19}$	$\text{C} \cdot \text{x}^{-1}$	
$\underline{q} / \underline{n}$	Faraday constant	$F = R / f = e \cdot N_A = 96\,485.332\,12$	$\text{C} \cdot \text{mol}^{-1}$	

\* The motive quantity with motive unit MU defines the physical format.

¶ Count  $N_X$  is the number  $N$  of elementary entities  $U_X$  with the abstract elementary unit [x] (Gnaiger 2020).

§ Redefinition of SI base units came into force on 2019-05-20; Bureau International des Poids et Mesures (2019) The International System of Units (SI), 9<sup>th</sup> edition.

§ A name or symbol was not found in the literature for the electromotive constant  $f$  introduced here.

¶ Elementary charge  $e \triangleq Q_{dp^+} / N_{p^+} = Q_{dp^+}$  is charge per proton count or charge per elementary proton  $U_{p^+}$ .

# Electrochemical Constants

**Jicheng Xie**



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