

# Outline of the Life of Christ

## Thirty Years Preparation

Mt

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Lk

Jn

Friday

1. Introductions.....
2. Birth Narratives.....
3. Magi, Egypt, Nazareth.....
4. Jesus in the temple at age 12.....
5. Baptized By John.....
6. Tempted by Satan.....
7. Jesus obtains his first disciples from John.....
8. First miracle: Water to wine at Cana.....

## First Year of Ministry

### Early Judean Ministry (8-9 months)

9. FIRST PASSOVER (AD 27), Jesus cleanses the temple.....
10. Interview with Nicodemus.....
11. Interview with the Woman at the well in Sychar of Samaria.....

### Galilean Ministry (1 year and 3 or 4 months)

12. From Cana, Jesus heals the royal official's son at Capernaum.....
13. Rejected at his home town Synagogue in Nazareth.....
14. Calls four fishermen.....
15. Heals the demoniac in the Capernaum Synagogue.....
16. Heals Peter's mother-in-law and many others in Capernaum.....
17. First Galilean Tour.....
  - (1) A leper healed.....
  - (2) Paralytic through the roof.....
  - (3) Call of Matthew; fight over feasting and fasting.....

## Second Year of Ministry

18. SECOND PASSOVER (AD 28)

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1  
 $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$   
1 mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and

titrimetric analysis the volume of a solution required to react

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pka of an unknown acid or pkb of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance  $\text{H}_2\text{C}_2\text{O}_4$  aq  $2\text{NaOH}$  aq  $\text{Na}_2\text{C}_2\text{O}_4$  4h20 to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction  $\text{NaOH}$  aq  $\text{C}_6\text{H}_4\text{COOH}$  aq  $\text{C}_6\text{H}_4\text{COONa}$  aq  $\text{H}_2\text{O}$  l the point of equivalence is the point of titration when the number of moles of naoh added is equal to the number of moles of khp in the

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