

1. Fundamentals of chemical kinetics

1.2 – Reaction order, rate constant, activation energy, half-life time

1.2.1 - Reaction order: m

For the reaction: $v_A A + v_B B \rightarrow \text{Products}$

$$v = k [A]^{m_A} [B]^{m_B} \quad (v \text{ expressed in } \text{mol m}^{-3} \text{ s}^{-1})$$

with k : rate constant ; m_A and m_B : partial orders

$$m_A + m_B = m: \text{ overall order }$$

If: $m_A = v_A$ and $m_B = v_B \iff$ ‘elementary reaction’

Fundamentals Of Chemical Kinetics

M Walker



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