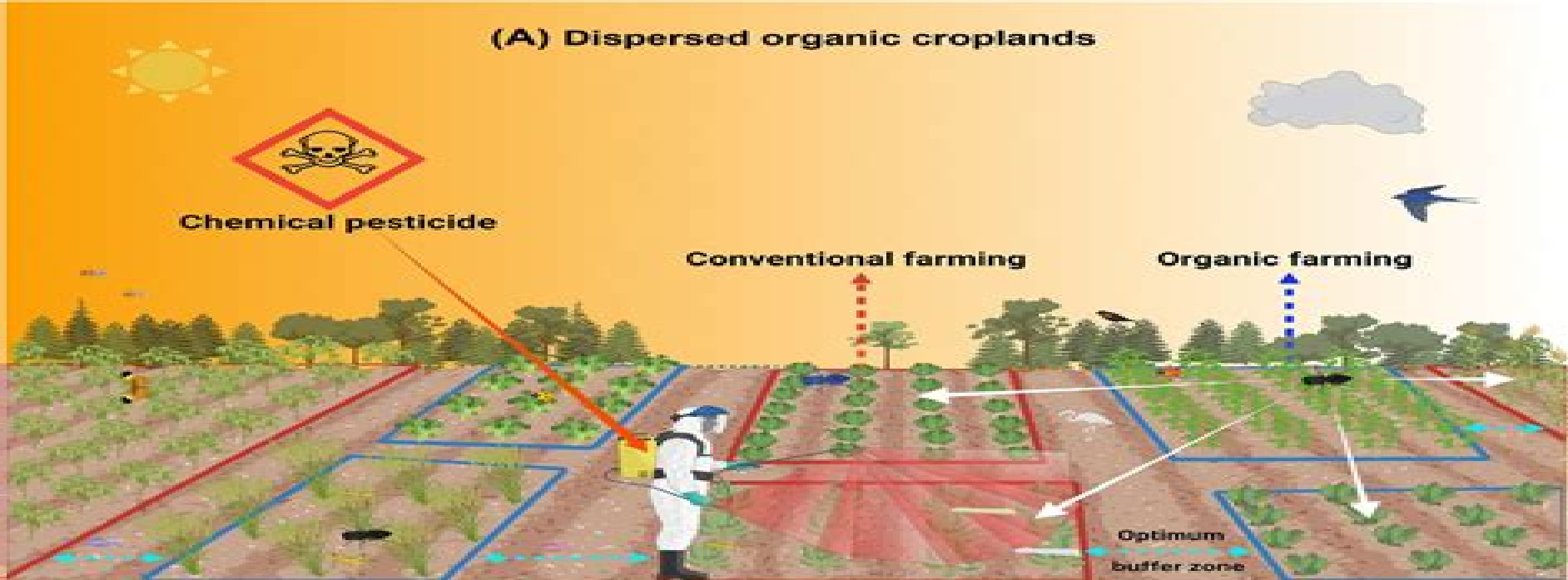
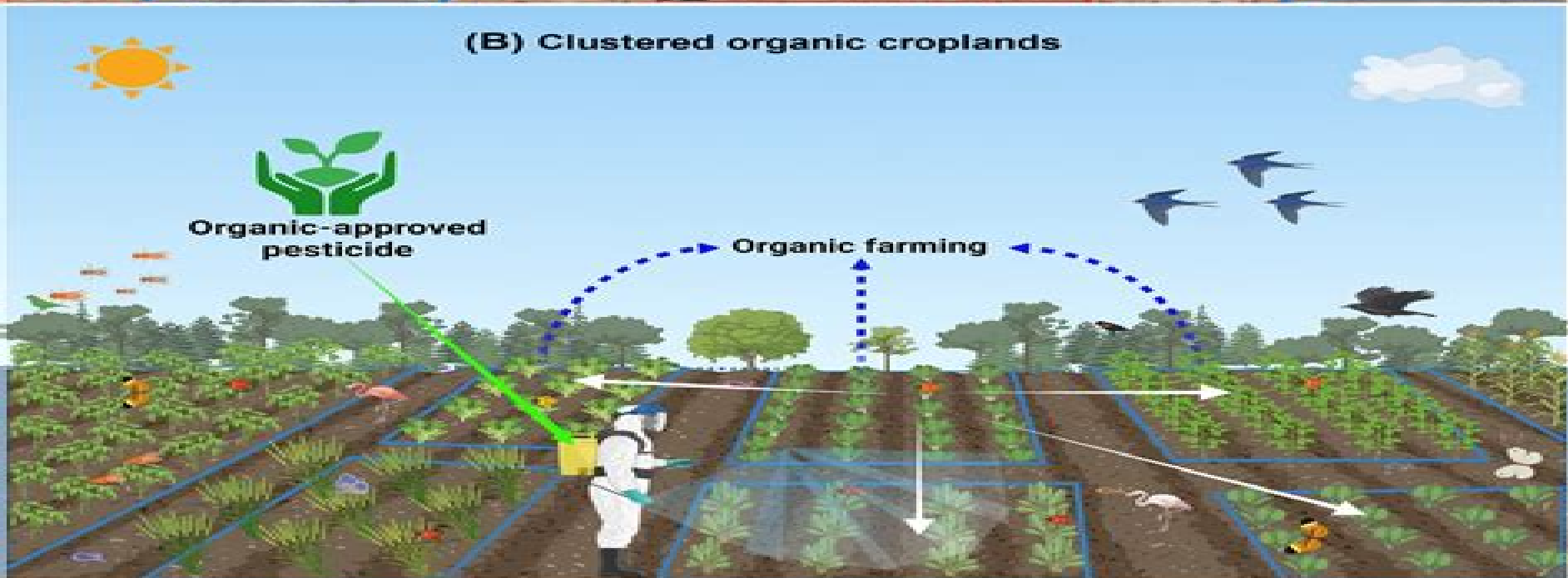


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# Ecology Of Pesticides

**DJ Losen**



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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} \rightarrow \text{Au}(\text{CN})_2^- + 2\text{e}^-$   $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  make electrons equal

convert to basic solution 4cn 2au 2au cn 2 2e multiplied by a factor of 2 2e 2h<sub>2</sub>O o<sub>2</sub> h<sub>2</sub>o<sub>2</sub> 2oh 3 the final answer

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web bonus problem here is a balanced net ionic equation 3br<sup>-</sup> 2 6oh<sup>-</sup> 5br<sup>-</sup> bro<sup>-</sup> 3 3h<sub>2</sub>O the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for

every atom of nickel oxidized but where do those electrons go

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