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net ionic equation $\text{H}^+\text{aq} + \text{HCO}_3^-\text{aq} \rightarrow \text{H}_2\text{O} + \text{CO}_2\text{g}$ here is a video on spectator ions
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web the six NO_3^-aq ions and the six Na^+aq ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+}\text{aq} + 2\text{PO}_4^{3-}\text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2\text{s}$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+\text{aq} + \text{OH}^-\text{aq} \rightarrow \text{H}_2\text{O l}$ the spectator ions in KOH aq and HNO_3aq are K^+aq and NO_3^-aq how to identify and list the

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loss or gain of one or more electrons ions containing positive charge are cations and those containing negative charge are anions an ion with a net charge of zero is electrically neutral

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be $2\text{cl aq} + \text{cu}^{2+} \text{aq} \rightarrow \text{cucl}_2 \text{ s}$ $2\text{cl aq} + \text{cu}^{2+} \text{aq} \rightarrow \text{cucl}_2 \text{ s}$ these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length