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the  $\text{OH}^-$  basic solution one in which the  $\text{H}^+$  is less than the  $\text{OH}^-$  alkaline solutions the  $\text{H}^+$  of a basic solution is less than  $1 \times 10^{-7} \text{ M}$

they are the same as basic solutions

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$\text{NH}_2$  and  $\text{S}^{2-}$  are  $\text{CH}_4$   $\text{NH}_3$  and  $\text{HS}^-$  respectively next we consider that acidity increases with positive charge on the molecule

thus ruling out that  $\text{S}^{2-}$  is the weakest base

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compounds have characteristics of both acids and bases in the same molecule we will explore the relationship between

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molecule that gains a hydrogen ion becomes this which is positively charged what did lewis propose an acid accepts a pair of

electrons during a reaction while a base donates a pair of electrons

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