

## 15.6: The effect of Temperature on Reaction rate

the rates of chemical reactions are, in general, highly sensitive to temperature.  $1^\circ \text{C} \approx 1.5$ ,  $10^\circ \text{C} \approx 2.5$

### The Arrhenius Equation

- the temperature dependence of the reaction rate is contained in the rate constant  $k$  (only constant when temp remains c).
- the modern form of the Arrhenius equation shows the relationship between the rate constant ( $k$ ) & the temp in  $k$  ( $T$ )

$$K = A e^{-\frac{E_a}{RT}}$$

FREQUENCY factor      exponential factor      activation  $E$

Activation energy ( $E_a$ ): an energy barrier or hump that must be surmounted in order for the reactants to transform into products.

Frequency factor ( $A$ ): the # of times that the reactants approach the activation barrier per unit time.



### THE ACTIVATION ENERGY

- to get from the reactant to the product, the molecule must go through a high-energy state called the activated complex. This is because energy is required to initially weaken the  $\text{H}_3\text{C}-\text{N}$  bond & allow the  $\text{NC}$  group to begin to rotate.

→ the energy required to reach the activated complex is the activation energy.

$\uparrow E_a \quad \downarrow$  reaction rate (at a given temp.)

### THE FREQUENCY FACTOR

- for the reaction, the frequency factor represents the rate at which the  $\text{NC}$  part of the molecule wags (wags side to side). With each wag, the reactant approaches the activation barrier.

↳ not equivalent to surmounting it.

# Essential Concepts Of Chemcustom

**SB Merriam**

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

## **Essential Concepts Of Chemcustom:**

## **Essential Concepts Of Chemcustom** Book Review: Unveiling the Magic of Language

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