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EDUCATION AND PRACTICE

**Upholding the Integrity of
Teaching and Learning**

Edited by Joseph Dunne and Pádraig Hogan

Education And Practice Upholding The Integrity Of Teaching And Learning

M Lipman



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the balanced equation shown below if the reaction of 90 6 grams of co produces 36 7 grams of C_3H_8 what is the percent
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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al $\times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 160.0 \text{ g}$ c determine the percent yield $\frac{101.96 \text{ g}}{160.0 \text{ g}} \times 100 = 63.7\%$

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 $80 \text{ g } I_2O_5 \times \frac{1 \text{ mol } I_2O_5}{283.5 \text{ g } I_2O_5} \times \frac{1 \text{ mol } I_2}{2 \text{ mol } I_2O_5} \times \frac{253.8 \text{ g } I_2}{1 \text{ mol } I_2} = 28 \text{ g } I_2$

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 $\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100$