



LWI  
THE DAYS OF  
LAURA INGALLS WILDER  
BOOK THREE

# G·O·O·D NEIGHBORS

A Novel by  
**T.L. TEDROW**

# Good Neighbors The Days Of Laura Ingalls Wilder Three

**Annelies Wilder-Smith**



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web answer  $\text{Na}^+ \text{aq}$  and  $\text{SO}_4^{2-} \text{aq}$  identify the chemical equation the spectator ions  $\text{KOH} \text{aq}$   $\text{HNO}_3$   $\text{KNO}_3 \text{aq}$   $\text{H}_2\text{O}$   $\text{K}^+ \text{aq}$   $\text{OH}^- \text{aq}$   $\text{H}^+ \text{aq}$   $\text{NO}_3^- \text{aq}$   $\text{K}^+ \text{aq}$   $\text{NO}_3^- \text{aq}$   $\text{H}_2\text{O}$  l answer  $\text{K}^+ \text{aq}$  and  $\text{NO}_3^- \text{aq}$  identify the chemical equation the spectator ions  $\text{Pb(NO}_3)_2 \text{aq}$   $2\text{KI} \text{aq}$   $\text{PbI}_2 \text{s}$   $2\text{KNO}_3 \text{aq}$

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction  $2\text{Ag}^+ \text{aq}$   $\text{Cr}_2\text{O}_7^{2-} \text{aq}$

rightarrow  $\text{Ag}^{2+} + 2\text{Cr}^{2+} + 7\text{S}^{2-}$  label 4 2 3

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web jan 29 2020 the ionic form of this reaction is  $2\text{Na}^{+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq}) + \text{Cu}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) \rightarrow 2\text{Na}^{+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + \text{CuCl}_2(\text{s})$  the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation  $\text{H}^{+}(\text{aq}) + \text{Cl}^{-}(\text{aq}) + \text{Na}^{+}(\text{aq}) + \text{HCO}_3^{-}(\text{aq}) \rightarrow \text{Na}^{+}(\text{aq}) + \text{Cl}^{-}(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$  the spectator ions are  $\text{Cl}^{-}$  and  $\text{Na}^{+}$  net ionic equation  $\text{H}^{+}(\text{aq}) + \text{HCO}_3^{-}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$  here is a video on spectator ions *chapter 8 04 ionic equations chemistry libretexts* - Mar 06 2023

web the six  $\text{NO}_3^{-}$  ions and the six  $\text{Na}^{+}$  ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation  $3\text{Ba}^{2+}(\text{aq}) + 2\text{PO}_4^{3-}(\text{aq}) \rightarrow \text{Ba}_3(\text{PO}_4)_2(\text{s})$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water  $\text{H}^{+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$

2 o l the spectator ions in koh aq and hno 3 aq are k aq and no 3 aq how to identify and list the

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